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The MoleThe Mole
280 Chemistry: Matter and Change • Chapter 14 Solutions Manual CHAPTER 14 SOLUTIONS MANUAL 32. Explain the similarities and differences between a 1 M solution of NaOH and a 1 m solution of NaOH. Both solutions contain NaOH (solute) dissolved in water (solvent). The 1 m solution contains 1 mole of NaOH per kilogram of water; the 1M solution 3 _ 5 _ _ _ 1 _ _ _ _ _ (

Mixtures and Solutions
Solutions Manual Chemistry: Matter and Change • Chapter 9 145 CHAPTER 9 SOLUTIONS MANUAL 35. Aqueous solutions of potassium iodide and silver nitrate are mixed, forming the precipitate silver iodide. chemical equation: KI(aq) AgNO 3(aq) 0 KNO 3(aq) AgI(s) complete ionic equation: K(aq) I (aq) 32 Ag (aq) NO 3 (aq) 0 K (aq) NO 3 (aq) AgI(s)

Chemical ReactionsChemical Reactions
86 Chemistry: Matter and Change • Chapter 6 Solutions Manual CHAPTER 6 SOLUTIONS MANUAL 4. Determine why producing 1 million fran-cium atoms per second is not enough to make measurements, such as density of boiling point. Even one million atoms collected together as a solid are microscopic. A grain of salt contains about 1015 sodium atoms.

The Periodic Table and Periodic LawThe Periodic Table and ...
124 Chemistry: Matter and Change • Chapter 8 Solutions Manual 45. O 3 O O O O O 46. Challenge Draw the Lewis resonance structure for the ion SO 3 2-. O S O 2- O OO O OO O S 2- S 2- Draw the expanded octet.Lewis structure of each molecule. 47. CIF 3 F Cl F F 48. PCI 5 Cl Cl Cl Cl P 49. Challenge Draw the Lewis structure for the molecule formed when six fluorine atoms and

Covalent BondingCovalent Bonding
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